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the nucleus and some electrons zipping around in orbitals when those pieces start combining in specific numbers
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devised and published by dmitri mendeleev in 1869 mendeleev found he could arrange the 65 elements then
known in a grid or table so that each element had 1 a higher atomic weight than the one on its left for example
magnesium atomic weight 24 3 is placed to the right of sodium atomic
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collection of element data the table lists the chemical elements in order of increasing atomic number which is the
number of protons in an atom of an element the rows periods and columns groups organize elements according
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name for its vertical columns each column is called a group the elements in each group have the same number of
electrons in the outer orbital those outer electrons are also called valence electrons they are the electrons

involved in chemical bonds with other elements

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easy to use chart of periodic table trends thoughtco Mar 12 2021 web 6 oct 2019 electronegativity electronegativity reflects how easily an atom can form a chemical bond generally electronegativity increases from left to right and decreases as you move down a group keep in mind the noble gasses column at the right hand side of the periodic table are relatively inert so their electronegativity approaches zero exception

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